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Chapter 12203. Chapter 12 Liquids:
Condensation, Evaporation, and
Dynamic Equilibrium. Review Skills 12.1
Change from Gas to Liquid and from
Liquid to Gas - An Introduction to

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Dynamic Equilibrium The Process of
Condensation The Process of
Evaporation How Evaporation Causes
Cooling Rate of Evaporation Dynamic
Equilibrium Between Liquid and Vapor
Equilibrium Vapor Pressure.

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Chemistry: Matter and Change Teacher
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Chapter 12 - States of Matter Section
12.1 Gases 1. motion 2. a. small b.
forces c. random d. elastic; kinetic 3. $KE = \frac{1}{2}mv^2$ 4. Temperature 5. true 6. true
7. false 8. true 9. true 10. false 11. true
12. false 13. a 14. a 15. d 16. d 17. b 18.

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Complete the table by writing the initial
and final phases for each phase change

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and making a check (✓) in the correct energy column. Phase Change Phase Energy released required initial final SO Column B boiling point freezing point melting point

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Chapter 12 - Chemical Kinetics |

CourseNotes

1. Chapter 12 - Chemical Kinetics. 12.1
Reaction Rates. A. Chemical kinetics 1.
Study of the speed with which reactants
are converted to products B. Reaction
Rate 1. The change in concentration of a
reactant or product per unit of time. t A t
 t concentration of A at time t

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concentration of A at time t Rate. 2 1 2
1.

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Matter and Change • Chapter 12 Section 12.2 Stoichiometric Calculations In your textbook, read about mole-to-mole conversion. Read the following passage

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and then solve the problems. In the equation that follows each problem, write in the space provided the mole ratio that can be used to solve the problem.

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