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Chemistry IF8766. Page 50.

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## **Chemistry If0235 Answers**

The molar mass of any substance is its atomic mass, molecular mass, or formula mass in grams per mole. The periodic table lists the atomic mass of carbon as 12.011 amu; the average molar mass of carbon—the mass of  $6.022 \times 10^{23}$  carbon atoms—is

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therefore 12.011 g/mol:

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When we say the molar mass of carbon



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is  $12.0 \text{ g mol}^{-1}$ , it means one mole of carbon weighs  $12.01 \text{ g}$ . In other words,  $6.022 \times 10^{23}$  atoms of carbon weigh  $12.01 \text{ g}$ . Important of Molar Mass. In chemistry, calculations are related to chemical reactions and stoichiometry.

**Molar Mass: Definition, Formula,  
Mole, Atomic Mass ...**

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mole = molecular weight / mass  
(multiply both sides by mass) mole \*  
mass = molecular weight (divide both  
sides by mole) mass = molecular weight  
/ mole. As  $1.626 \times 10^{23}$  molecules of  
NaOH is also equal to 0.27 moles, and  
we know that the molecular weight of  
NaOH is 40, we can use these numbers  
to get: mass =  $40 / 0.27 = 10.8$  g

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## **Mole Calculator**

Conversions like this are possible for any substance, as long as the proper atomic mass, formula mass, or molar mass is known (or can be determined) and expressed in grams per mole. Figure  $\{\text{PageIndex}\{1\}\}$  is a chart for determining what conversion factor is

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needed, and Figure  $\{\{PageIndex\{2\}\}$  is a flow diagram for the steps needed to perform ...

## **5.4: Molar Mass- Mole-to-Mass and ... - Chemistry LibreTexts**

Describes calculations involving conversions between gas volume, number of moles, and mass of the gas.

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## **Conversions between Moles and Gas Volume ( Read ...**

Number of moles of NaOH = mass  $\div$   
relative formula mass =  $20 \div 40 = 0.5$   
mol From the equation, 2 mol of NaOH  
reacts with 1 mol of Na<sub>2</sub>SO<sub>4</sub>, so 0.5  
mol of NaOH will react with 0.25 mol of  
Na<sub>2</sub> ...

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## **Mole calculations - Formula mass and mole calculations ...**

One mole of iron has a mass of 56 g.

Moles of compounds A mole of a  
molecular compound contains  $6 \times 10^{23}$   
molecules. It has a mass that is equal to  
its relative formula mass.

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## **The mole - Formula mass and mole calculations - GCSE ...**

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One mole of a substance is equal to  $6.022 \times 10^{23}$  units of that substance



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(such as atoms, molecules, or ions). The number  $6.022 \times 10^{23}$  is known as Avogadro's number or Avogadro's constant. The concept of the mole can be used to convert between mass and number of particles.

**The mole and Avogadro's number  
(video) | Khan Academy**

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The units of mass are typically grams. Mass percent is also known as percent by weight or w/w%. The molar mass is the sum of the masses of all the atoms in one mole of the compound. The sum of all the mass percentages should add up to 100%. Watch for rounding errors in the last significant figure to make sure all the percentages add up.

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'Chemistry 801 Mole Mole and Mole

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## **Chemistry Lab Stoichiometry Mole**

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## **And Mass Relationships**

Molar Mass; In order to differentiate them (as well as see how they are connected to each other), we have to know their definitions first. Let's take a look at them one at a time..... A)

Relative Atomic Mass,  $A_r$ . The relative atomic mass of an atom is the average mass of one atom of that element

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compared to  $1/12$  of the mass of one carbon-12 atom.

### **Molar Mass - O Level Chemistry & IP Chemistry Notes by 10 ...**

In chemistry, the molar mass of a chemical compound is defined as the mass of a sample of that compound divided by the amount of substance in



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that sample, measured in moles. The molar mass is a bulk, not molecular, property of a substance. The molar mass is an average of many instances of the compound, which often vary in mass due to the presence of isotopes.

### **Molar mass - Wikipedia**

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Density of water is  $0.99893 \text{ g cm}^{-3}$ , and the molar mass of air is  $28.8 \text{ g mol}^{-1}$ . Find the molar mass of carbon dioxide. I have calculated the mass of  $\text{CO}_2$ ,

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using the mass of the flask filled with gas and subtracting the mass of the empty flask

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